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neutralization happens because hydrogen ions and hydroxide ions react to produce water  $\text{H}^+ + \text{OH}^- \rightarrow \text{H}_2\text{O}$   
h 2 o l hydrochloric

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 $\text{acid}_1 + \text{base}_2 \rightleftharpoons \text{acid}_2 + \text{base}_1$

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web the  $\text{H}^+$  ion is a base because it is accepting a proton from the  $\text{H}_3\text{O}^+$  the reversible reaction contains two acids and two bases we think of them in pairs called conjugate pairs when the acid  $\text{H}_3\text{O}^+$  loses a proton it forms a base  $\text{H}_2\text{O}$  when the base  $\text{H}_2\text{O}$  accepts a proton back again it obviously reforms the acid  $\text{H}_3\text{O}^+$

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