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titanium tetrachloride  $\text{TiCl}_4$  has a

melting point of  $23.2^\circ\text{C}$  and has a  $\Delta h$

fusion  $9.37\text{ kJ/mol}$  how much energy is required to melt  $263.1\text{ g TiCl}_4$  for  $\text{TiCl}_4$

which will likely have the larger magnitude  $\Delta h$  fusion or  $\Delta h$  vaporization explain your reasoning answer a

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graphs they plot a substance's

temperature y axis against heat x axis

for heating curves we start with a solid

and add heat energy for cooling curves

we start with the gas phase and remove

heat energy cooling and heating curves

have five segments let's study each one

by looking at heating ice to

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multiple choice 2 minutes 1 pt describe

the substance between letters d and e

gas liquid melting evaporating 2

multiple choice 2 minutes 1 pt between

which points is the temperature of the

substance remaining constant a b only

a b c d e f b c only b c d e 3 multiple

choice 2 minutes 1 pt

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