

The Prepared Nurse Co

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342g of 20 by mass of b a ba oh 2 solution sp gr 0 57 is
reacted with 1200 ml of 2m hno3 if the final density of
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solution is 1 0 25m 2 0 5m 3 0 888m 4 none of

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solution of hcl is 1 0585 g ml the molality of the solution is 1
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molarity is in the denominator molality is measured in
kilograms of solvent whereas molarity is measured in liters
of solution when studying the properties of solutions related
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signifies solutions labelled with molar concentration a 1 0 m

solution contains 1 mole of solvent per litre of solution

molality is the number of solvent moles per kilogram

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00581158 mol 0 00581158 mol 0 4500 kg 0 0129 m 3 mole

fraction na 2 hpo 4 0 825 g 141 9579 g mol 0 00581158 mol

h 2 o 450 0 g 18 015 g mol 24 97918401 mol

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4 2 2 a first convert 250 ml to liters 250 1000 0 25 then

calculate molarity 5 moles 0 25 liters 20 m 3 c a solution

with molarity 2 requires 2 m of n a oh per liter so 4 x 2 8 m 4

a a solution of molarity 1 5 m requires 1 5 mol of na to every

litre of solvent

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of a mixture mmix can be calculated using the following

formula $m_{mix} = \frac{m_1}{V_1} = \frac{m_2}{V_2} = \frac{V_1}{V_2} = \frac{1.5 \times 480}{1.2 \times 520}$

$480 \times 520 = 720 \times 624 = 1000 \times 344$ m q5 with increase in

temperature which of these changes a molality b fraction of

solute present in water

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40 mol of a substance dissolved in 1 6 l of a solution 2 what

is the molarity of a solution containing 325 g of nacl

dissolved in 750 ml of solution 3 140 g of kcl is dissolved in

600 ml of water what is the molarity 4 724 4 g of ammonium

phosphate in 4500 ml of alcohol

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solutions 0 444 mol of cocl 2 in 0 654 l of solution 98 0 g of

phosphoric acid h 3 po 4 in 1 00 l of solution 0 2074 g of

calcium hydroxide ca oh 2 in 40 00 ml of solution 10 5 kg

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while molarity is based on the liters of solution molality is

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