

# 10 Sentences About Karachi

JR Anderson

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web this data shows that there was a gain the mass of the lighter which

shows that there was a certain amount of butane 0 38 grams 0 02 grams gas that was successfully collected during the gas collection there is a mild influence of barometric pressure during both trials as shown above in the raw data

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web 11 use avogadro s law to determine the number of moles of butane gas assume that butane is an ideal gas and that one mole has a volume of 22 4 l at stp v 1 n 1 v 2 n 2 0 0896 l n 22 4 l 1 mole n 0 0040 moles 12 experimental molar mass g mole of butane 0 23 g 0 0040 moles 58 g mole 13 the molecular formula of butane is c 4 h 10

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web nov 3 2020 in this experiment the moles and mass will be required to determine the molar mass of butane in a butane lighter experimental to begin the experiment the mass of a butane lighter was measured before using the butane gas to deplete the water to the 80ml mark on a graduated cylinder

**classroom resources determination**

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web the pressure of the gas the pressure of the gas was calculated using dalton s law of partial pressures by subtracting the pressure of the water vapor from the total pressure

which in this case was the atmospheric pressure 760mm hg the pressure of the water vapor was found by using the temperature of the water

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web jun 18 2019 an ideal gas follows the ideal gas law at all conditions of p and t the particles in an ideal gas do not have finite size and volume the collisions between the ideal gas particles are said to be elastic they exert no attractive or repulsive forces hydrogen gas generated in today s experiment is however a real gas not an ideal gas

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web in this lab you will collect a measurable volume of butane gas using water displacement using dalton's law and the ideal gas law the number of moles of butane gas and then the molar mass of butane can be calculated your experimental value for molar mass will be compared to the known molar mass of butane and percent error calculated

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web use the combined gas law to determine the volume in l of butane at stp use avogadro's law to determine the number of moles of butane gas assume

that butane is an ideal gas and that one mole has a volume of 22 l at stp experimental molar mass g mole of butane the molecular formula of butane is c 4 h 10

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web although the most convenient gas for this is butane other gases may be available in similar small cylinders the collection of a gas sample and the weighing of the gas cylinder before and after this should take about 5 10 minutes as a demonstration

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web she later earned an ma in english from colorado state university where she taught for two years and an mfa in writing for children and young adults from vermont college

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