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Brendan G. Carr

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carbon monoxide co determine the mass of iodine I_2 which could be produced b if in the above situation only 0.160 moles of iodine I_2 was produced i what mass of iodine was produced

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web start with 8.3 mole H_2S and multiply it with 2 mole H_2O over 2 mole H_2S from the equation $8.3 \text{ mole } H_2S \times \frac{2 \text{ mole } H_2O}{2 \text{ mole } H_2S} \times 18 \text{g } H_2O = 149.4 \text{g } H_2O$ is your theoretical yield so percent yield $\frac{137.1 \text{g } H_2O}{149.4 \text{g } H_2O} \times 100 = 91.77\%$ problems 6 8 are similar to problem 5

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experimental design yield efficiency mass of product
obtained calculated mass of the product expected $\times 100$ in
this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$
4 $\text{Al} + 3 \text{O}_2 \rightarrow 2 \text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3
160.0 g $\text{Al} \times 1 \text{ mol Al} \times 2 \text{ mol Al}_2\text{O}_3 \times 101.964 \text{ g Al}_2\text{O}_3 / 302.3 \text{ g Al}_2\text{O}_3$
26.982 g $\text{Al} \times 4 \text{ mol Al} / 1 \text{ mol Al}_2\text{O}_3$ c determine the
percent yield $\text{yield actual} \times 100 / 260.0 \text{ g} \times 100 / 86.01$
theoretical 302.3 g

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web 7 Na_2CO_4 is the limiting reactant percent yield 86.6
8 only four molecules can be made 9 this amount cannot be
weighted by ordinary balances and is worthless 10 nitrogen
is the limiting reagent 11 yes methane is the limiting reagent
12 c is the limiting reagent 4.33 g of H_2 are left over 13

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reagent from a set of reactants calculate how much product
will be produced from the limiting reagent calculate how
much reactant s remains when the reaction is complete

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web the percent yield is calculated as follows text percent
yield $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100$
percent yield theoretical yield actual yield 100 based on this

definition we would expect a percent yield

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reactant limiting reagent in a given chemical reaction
calculate how much product will be produced from the
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methanol CH_3OH can be produced by the following reaction
 $2\text{H}_2 + \text{CO} \rightarrow \text{CH}_3\text{OH}$ calculate the theoretical yield of CH_3OH if 68.5
g of CO is reacted with 8.6 g of H_2

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stoichiometric problems from a balanced chemical equation
3 sto 5 differentiate between the actual yield and theoretical
yield of a chemical reaction

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consider the following reaction $\text{NH}_4\text{NO}_3 + \text{Na}_3\text{PO}_4 \rightarrow \text{NH}_4\text{}_3\text{PO}_4$
 NaNO_3 which reactant is limiting assuming we started with
30.0 grams of ammonium nitrate and 50.0 grams of sodium
phosphate what is the

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consider the reaction I_2O_5 g 5 co g 5 CO_2 g i2 g a 80 0 grams
of iodine v oxide I_2O_5 reacts with 28 0 grams of carbon
monoxide co determine the mass of iodine i2 which could be
produced 80 g I_2O_5 1 mol I_2O_5 1 mol i2 xs 1 333 8 g I_2O_5 1
mol I_2O_5 28 g co 1 mol co

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web jun 2 2020 the percent yield of a reaction is the ratio of
the actual yield to the theoretical yield multiplied by 100 to
give a percentage text percent yield text actual yield g over
text theoretical yield g times 100 label 3 7 3

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university was established 780 acres the atmosphere is quite
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