

Synopsis Of People We Meet On Vacation

Jin-Ying Zhang

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containing 12 or more carbon atoms per molecule the essential properties of kerosene are flash point distillation range burning characteristics sulfur content color and cloud point acids can be present in kerosene aviation turbine fuels due to acid treatment during refining
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nfpa health rating

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web allow thegaseous phase thermodynamics to be curve fittedto 300k about142k lowerthanthe normalboiling point 542k 2of kerosene nevertheless in actualcalculationsinvolving liquid kerosene fuel the heatcapacityandlatentheatof vaporizationof liquid kerosene 12areusedfor

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web as a general trend vapor pressures of liquids at ambient temperatures increase with decreasing boiling points this is illustrated in the vapor pressure chart see right that shows graphs of the

vapor pressures versus temperatures for a variety of liquids

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13 10 vapor pressure curves

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web the graph of the vapor pressure of water versus temperature in figure pageindex 3 indicates that the vapor pressure of water is 68 kpa at about 90 c thus at about 90 c the vapor pressure of water will equal the atmospheric pressure in

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 expresses the volume occupied by 1
 mole of that respective gas under
 certain temperature and pressure
 conditions the most common example is
 the molar volume of a gas at stp
 standard temperature and pressure
 which is equal to 22.4 l for 1 mole of
 any ideal gas at a temperature equal to
 273.15 K and a pressure

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web the molar volume will be the
 volume occupied by one molar mass 32
 g of the liquid expressing the density in
 liters instead of ml we have $V_m = \frac{m}{\rho}$
 $\frac{32 \text{ g}}{1.790 \text{ g/L}} = 17.87 \text{ L/mol}$
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 Avogadro's law the volume is directly
 proportional to the number of moles the
 important point to keep in mind is that
 n can be moles of any gas since one
 mole always contain 6.02 $\times 10^{23}$

molecules and one molecule of any gas
 at the same temperature will have the
 same kinetic energy based on this
 principle it is found
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 web apr 22 2020 molar volume of
 gases 1 counting in moles one mole is
 defined as 6.02 $\times 10^{23}$ particles this is the
 Avogadro number an unchanging
 constant whether you are counting
 hydrogen atoms carbon dioxide
 molecules or carbonate ions a mole is
 always 6.02 $\times 10^{23}$ but as different
 particles have different mass their
 molar mass varies

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 concentration of a gas though and that
 calculation would be exactly the one
 you've described placing x moles of a
 gas in a container of volume V yields a
 concentration $C = \frac{x}{V}$ with units of e.g. mol/L

mol over dm³

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web the molar volume has the SI unit of
 cubic metres per mole m³ mol
 although it is more typical to use the
 units cubic decimetres per mole dm³
 mol for gases and cubic centimetres
 per mole cm³ mol for liquids and solids
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web zero degrees Celsius or 273 K is
 defined as standard temperature and
 one atmosphere or 760 mm Hg is
 defined as standard pressure the
 volume occupied by one mole of a gas is
 called molar volume therefore the
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 chemical element or a chemical
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 dividing the molar mass M by mass
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properties of gases such as

temperature pressure and volume are

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finally the collisional behavior of gases

is used to

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volume and from gas volume to moles

the equality of 1 mol 22 4 l is the basis

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note that for big values hundreds of

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real gases divert from ideal gas law

that s why they are not ideal and this

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be calculated from its number of moles

using volume of gas moles vm

rearranging this moles of gas frac

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latm molk 273k 1 00atm v 22 4l in

other words 1 mole of a gas will occupy

22 4 l at stp assuming ideal gas

behavior at stp the volume of a gas is

only dependent on number of moles of

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liquids are measured in cubic

centimetres cm³ or cubic decimetres

dm³ or litres l 1 mole of any gas at

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