

# Stuck In Neutral Summary

S Baum

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dry thoroughly with a paper towel 2  
place the penny on a fresh paper towel  
3 fill a beaker with 25 ml of tap water 4  
using a pipet slowly drop individual  
droplets of water onto the surface of  
the penny 5 count each drop until the  
water begins to spill over the sides of  
the penny

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questions be sure to complete your data  
table including observations from the  
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water they think a coin can hold were  
your predictions correct does it matter  
if the coin is heads or tails will syrup oil  
hold more or less drops than water why  
how is it possible to get so many water  
drops on a coin what causes the skin on  
the surface of the big droplet

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alloy a mixture of 2 or more metals or a  
metal and a nonmetal fused together  
dissolved in each other when molten  
low zinc brasses contain how much zinc  
and how do they form less than 20 zinc  
easier to form how much zinc do yellow  
brasses contain 34 37 zinc what other  
elements do some brasses contain

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known as surface tension we will  
observe and study this property by  
determining how many drops of water  
will fit on

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tension of water affected by soap  
introduction surface tension refers to  
water s ability to stick to itself surface  
tension can be measured and observed  
by dropping water drop by drop onto a  
penny the number of water drops that  
can fit on a penny will surprise you 1

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answer for this lab is w s the percent  
copper of a post 1983 1959 to 1982  
copper made up 95 of pennies while ti n  
and zinc made up the remaining 5  
however since copper became m ore  
expensive the united states treasury  
took t sion to alter the pennies

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web in this activity you ll experience  
surface tension and cohesion by testing  
how many drops of water you can you  
fit on a penny you might be surprised  
considering pennies are so small there  
s only one way to find out

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web what was the purpose of the penny  
lab observe how properties of a metal

can change how will copper be changed in terms of its properties by heating the copper with zinc in the presence of a sodium hydroxide naoh solution the zinc replaces the copper in the naoh solution then the zinc diffuses into the copper when heated zinc naoh

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web 1 explain your results from both parts of the experiment in terms of cohesion and surface tension 2 how do your results compare to the other groups in your class provide at least 2 possible reasons for any similarities and differences you identified drops on a penny teacher notes materials

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