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**Didier Musso**

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moles of butane in order to find the molarity the mass of the  
butane should be divided by the moles after further  
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be 56 81 grams moles this lab is done with tons of sources of  
error

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web this data shows that there was a gain the mass of the  
lighter which shows that there was a certain amount of  
butane 0 38 grams 0 02 grams gas that was successfully  
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of barometric pressure during both trials as shown above in  
the raw data

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n 22 4 l 1 mole n 0 0040 moles 12 experimental molar mass g  
mole of butane 0 23 g 0 0040 moles 58 g mole 13 the  
molecular formula of butane is c 4 h 10

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butane lighter experimental to begin the experiment the  
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