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 reactions and identify each type of
 reaction NaBr Ca(OH)_2 CaBr_2 NaOH NH_3
 H_2SO_4 NH_4_2SO_4 $\text{C}_5\text{H}_9\text{O}$ O_2 CO_2 H_2O Pb
 H_3PO_4 $\text{H}_2\text{Pb}_3\text{PO}_4$ $2\text{Li}_3\text{N}$ NH_4NO_3 LiNO_3
 NH_4NO_3

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 a red b blue c green d yellow 2 use
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 the equations to help you determine the
 type of reaction it illustrates use the

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 reaction $2\text{H}_2\text{O}$ 2H_2 O_2 is a synthesis
 reaction decomposition reaction single
 displacement reaction double
 displacement reaction combustion
 reaction you can recognize a
 decomposition or analysis reaction by
 noticing a compound or molecule
 breaks into smaller chemical species 2
 the chemical reaction

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 concerned with understanding the rates
 of chemical reactions it is to be
 contrasted with thermodynamics which
 deals with the direction in which a
 process occurs but in

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 organisation of the known chemical
 elements from which they can make
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 1 contain the phase notations s l g and
 aq write each symbol by each meaning
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 g 2 based on the

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1 $\text{Na}_3\text{PO}_4 + 3\text{KOH} \rightarrow 3\text{NaOH} + \text{K}_3\text{PO}_4$ double displacement

2 $\text{MgCl}_2 + \text{Li}_2\text{CO}_3 \rightarrow \text{MgCO}_3 + 2\text{LiCl}$ double displacement

3 $\text{C}_6\text{H}_{12} + 9\text{O}_2 \rightarrow 6\text{CO}_2 + 6\text{H}_2\text{O}$ combustion

4 $\text{Pb} + \text{FeSO}_4 \rightarrow \text{PbSO}_4 + \text{Fe}$ single displacement

5 $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$ decomposition

6 $\text{P}_4 + 3\text{O}_2 \rightarrow 2\text{P}_2\text{O}_3$

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1 ch $14\text{O}_2 + 10\text{C}_9\text{H}_{20} \rightarrow 2\text{C}_9\text{H}_{18}\text{O} + 3\text{H}_2\text{O}$

2 $\text{AgCl} + 3\text{LiBr} \rightarrow 2\text{LiF} + \text{Br}_2 + 4\text{NaF} + 5\text{F}_2$

2 $\text{Br}_2 + \text{B}$ balance the chemical reactions and identify each type

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web 1 $3\text{NaBr} + \text{H}_3\text{PO}_4 \rightarrow \text{Na}_3\text{PO}_4 + 3\text{HBr}$ type of reaction double displacement

2 $3\text{CaOH} + 2\text{Al}_2\text{SO}_4 \rightarrow 3\text{CaSO}_4 + 2\text{AlOH}$ type of reaction double displacement

3 $3\text{Mg} + \text{Fe}_2\text{O}_3 \rightarrow 2\text{Fe} + 3\text{MgO}$ type of reaction single displacement

4 $\text{C}_2\text{H}_4 + 3\text{O}_2 \rightarrow 2\text{CO}_2 + 2\text{H}_2\text{O}$

type of reaction combustion 5

2pbso42pbso3 1o2 type of reaction

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