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including spectator ions are explicitly represented by formulas for their dissociated ions molecular equation chemical equation in which all reactants and products are **spectator ion definition and examples thoughtco** - Nov 02 2022 web jan 29 2020 the ionic form of this reaction is 2 na aq 2 cl aq cu 2 aq so 4<sup>2-</sup> aq 2 na aq so 4<sup>2-</sup> aq cucl 2 s the sodium ions and sulfate ion are the spectator ions in this reaction they appear unchanged in both the product and reactant side of the equation

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spectator ions an ion that is not directly involved in a chemical reaction an ion that does not change oxidation number or composition during a reaction  
products

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web the equation below is the final net ionic equation after elimination of the spectator ions which yields water  
 $\text{h}^+ \text{aq} + \text{oh}^- \text{aq} \rightarrow \text{h}_2\text{O} \text{l}$   
the spectator ions in  $\text{koh} \text{aq}$  and  $\text{hno}_3 \text{aq}$  are  $\text{k}^+ \text{aq}$  and  $\text{no}_3^- \text{aq}$   
how to identify and list the

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 $\text{pb}(\text{no}_3)_2 \text{aq} + 2\text{kl} \text{aq} \rightarrow \text{pb}(\text{s}) + 2\text{kno}_3 \text{aq}$   
complete ionic equation an equation written with all soluble strong electrolytes shown as ions

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web the  $\text{ag}^+ \text{aq}$  and  $\text{cl}^- \text{aq}$  ions become  $\text{agcl} \text{s}$  but the  $\text{na}^+ \text{aq}$  ions and the  $\text{no}_3^- \text{aq}$  ions stay as  $\text{na}^+ \text{aq}$  ions and  $\text{no}_3^- \text{aq}$  ions these two ions are examples of spectator ions ions that do nothing in the overall course of a chemical

reaction

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web aug 8 2022 a spectator ion is an ion that does not take part in the chemical reaction and is found in solution both before and after the reaction the net ionic equation is the chemical equation that shows only those elements compounds and ions that are directly involved in the chemical reaction

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of the reactions  
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 web jun 2 2023 the spectator ions are cut out from a reaction hence the net ionic equation for the above reaction

would be  $2\text{Cl}^-\text{aq} + \text{Cu}^{2+}\text{aq} + 2\text{CuCl}_2\text{s} + 2\text{Cl}^-\text{aq} + \text{Cu}^{2+}\text{aq} + 2\text{CuCl}_2\text{s}$  these spectator ions are not actively involved in the reaction the existence of spectator ions affects the debye length