

# Gilles Deleuze The Fold Leibniz And The Baroque

Wenbin Ji

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$$\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu(NO}_3)_2 + 2\text{Ag}$$
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$$\text{C}_3\text{H}_8 + 5\text{O}_2 \rightarrow 3\text{CO}_2 + 4\text{H}_2\text{O}$$
 7h 2 c 3 h 8 3h 2 o  
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 web start with 8 g mole h<sub>2</sub>s and multiply it with 2 mole h<sub>2</sub>o over 2 mole h<sub>2</sub>s from the equation  

$$8 \text{ g h}_2\text{S} \times \frac{2 \text{ mole h}_2\text{O}}{2 \text{ mole h}_2\text{S}} \times \frac{18 \text{ g h}_2\text{O}}{1 \text{ mole h}_2\text{O}} = 144 \text{ g h}_2\text{O}$$
 149 4g h<sub>2</sub>o is your theoretical yield so percent yield  

$$\frac{149 \text{ g}}{144 \text{ g}} \times 100 = 103.4\%$$
 149 4g h<sub>2</sub>o x 100 91 77 problems 6 8 are similar to problem 5  
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 web 1 csun edu hcchm001 introchemhandouts html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of nh<sub>3</sub> can be produced theoretically from the reaction of 50 g of n  

$$\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$$
 what is the limiting reagent if 85 g are actually formed what is the  
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 web limiting reagents problem one 1 for the balanced equation  

$$\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO}_2 + 3\text{H}_2\text{O}$$
 if the reaction of 21 g of o<sub>2</sub> produces a 47 g yield how many grams of co would be produced step one figure out how many moles of the limiting reagent you have  

$$\frac{21 \text{ g}}{32 \text{ g/mol}} = 0.656 \text{ mol}$$
 21 g 32 g mole 21 g 0 66 moles 32 g m  
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 web percent yield is a measure of the efficiency of the experimental design  

$$\text{Percent Yield} = \frac{\text{mass of product obtained}}{\text{calculated mass of the product expected}} \times 100$$
 in this reaction a yield 80 is anticipated  
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 web a write a balanced equation for the reaction  

$$\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$$
 4 al 3 o2 2 al2o3 b determine the theoretical yield of al<sub>2</sub>o<sub>3</sub>  

$$160 \text{ g Al} \times \frac{1 \text{ mol Al}}{27 \text{ g Al}} \times \frac{2 \text{ mol Al}_2\text{O}_3}{4 \text{ mol Al}} \times \frac{101 \text{ g Al}_2\text{O}_3}{1 \text{ mol Al}_2\text{O}_3} = 302 \text{ g Al}_2\text{O}_3$$
 160 0 g al x 1 mol al x 2 mol al2o3 x 101 964 g al2o3 302 3 g al2o3 26 982 g al 4 mol al 1 mol al2o3 c determine the percent yield  

$$\frac{260 \text{ g}}{302 \text{ g}} \times 100 = 86.1\%$$
 yield actual x 100 260 0 g x 100 86 01 theoretical 302 3 g  
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 web 7 na 2 c 2 o 4 is the limiting reactant percent yield 86 6 8 only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 c is the limiting reagent 4 33 g of h 2 are left over 13  
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web the percent yield is calculated as follows text percent yield  $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100$  percent yield theoretical yield actual yield 100 based on this definition we would expect a percent yield

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practice worksheet answer key limiting reactant yield practice worksheet 1 methanol  $\text{CH}_3\text{OH}$  can be produced by the following reaction  $2\text{H}_2 + \text{CO} \rightarrow \text{CH}_3\text{OH}$  calculate the theoretical yield of  $\text{CH}_3\text{OH}$  if 68.5 g of CO is reacted with 8.6 g of  $\text{H}_2$

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