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JA Banks

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web jan 26 2016 the solubility product or ksp is a way of showing how effective a chemical is at dissolving into an aqueous solution the higher the ksp the more the chemical will dissolve into that solvent for this lab calcium hydroxide is created then ionized Ca(OH)_2 when calcium is put in water reacts as follows $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$

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web k sp of calcium hydroxide 1 author j m mccormick last update ocober 13 2013 introduction an equilibrium constant k is related to ΔG for a process through eqn 1 and as long as ΔH and ΔS are independent of temperature eqn 1 can be combined with eqn 2 to determine ΔH and ΔS from the temperature dependence of k

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web in today s experiment you will determine the solubility product ksp of calcium hydroxide Ca(OH)_2 by measuring the concentration of Ca(OH)_2 in a saturated solution calcium hydroxide is a sparingly soluble salt that dissolves according to the following reaction the solubility product expression for this reaction is

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web expert verified chem 182 determining the ksp of calcium hydroxide calcium hydroxide is an ionic solid that is sparingly soluble in water a saturated aqueous solution of Ca(OH)_2 is represented in equation form as shown below $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$ the solubility product expression describes in mathematical terms the equilibrium

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web ksp for calcium hydroxide calcium hydroxide is a soft white caustic powder used in making mortar cements calcium salts paints and petrochemicals it is also used in saltwater aquaria to make up kalkwasser limewater solutions for reef tanks and is used as a ph regulating agent notice that calcium hydroxide is divalent and thus has twice *ksp calcium hydroxide chem132 determination of the ksp of calcium* - Jun 24 2022

web use the concentration of calcium hydroxide to calculate the concentration of calcium ions and hydroxide ions in the solution $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$ $5 \times 10^{-2} \text{ M}$ calcium ions $5 \times 10^{-2} \text{ M}$ hydroxide ions calculate calcium hydroxide ksp using the concentrations of calcium ions and hydroxide ions

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the objective of this lab was to measure the ksp of a slightly
soluble basic salt and how solubility changes with
temperature ksp is another form of equilibrium expression
that deals with solubility the reaction that will take place is
 $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$

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hydroxide post lab report sheet data table 2 points trial 1
trial 2 hci molarity 0.050 m volume of filtered ca oh 2
solution ml 5.00 5.00 3.71 13.75 equivalence point volume ml
oh in saturated ca oh 2 m ca in saturated ca oh 2 m ksp of ca
oh 2 average ksp of ca oh 2

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web the chemical equation used was $\text{Ca(OH)}_2 \text{ (s)} + 2\text{HCl (aq)} \rightarrow \text{CaCl}_2 \text{ (aq)} + 2\text{H}_2\text{O (l)}$
the solubility product was determined to be $K_{\text{sp}} \text{ Ca(OH)}_2$ from
the titration which was made by dripping a concentrated
solution of calcium hydroxide into hydrochloric acid the
calcium hydroxide was filtered before starting the titration
18.1 solubility product constant K_{sp} chemistry libretexts -
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web jul 12 2023 because the concentration of a pure solid such as $\text{Ca}_3(\text{PO}_4)_2$ is a constant it does not appear explicitly in the equilibrium constant expression the equilibrium constant expression for the dissolution of calcium phosphate is therefore $K = \frac{[\text{Ca}^{2+}]^3 [\text{PO}_4^{3-}]^2}{[\text{Ca}_3(\text{PO}_4)_2]}$ $K_{\text{sp}} = [\text{Ca}^{2+}]^3 [\text{PO}_4^{3-}]^2$

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web lab report 8 determinin of calcium hydroxide the purpose of this was to titrat e a saturated Ca(OH)_2 soultion with a st andard hcl solution and find the ksp or the oh concentration was f ound and then the concentration was found and u for this experiment a titration of calcium hydroxide with hydrochlo

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web therefore it was determined through error calculations and ease of use that the most reliable valid method to use to get the closest ksp value to the theoretical ksp for ca was the determination through ph method 2 introduction calcium hydroxide is a strong base soluble in water

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