

Introduction To The Thermodynamics Of Materials

David R Gaskell

Julia Schneider

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 $q_{\text{calorimeter}} = m \times c \times \Delta t$ 7.3.14 7.3.14 $\Delta H_{\text{rxn}} = q_{\text{rxn}}$ $q_{\text{calorimeter}} = m \times c \times \Delta t$

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water calculate the molar heat of combustion of hexane if the water temperature rises 55.4 $^\circ\text{C}$ the specific heat capacity of water is 4.184 J $\text{g}^{-1} \text{ } ^\circ\text{C}^{-1}$ 5.650 grams H_2O 4.184 J $\text{g}^{-1} \text{ } ^\circ\text{C}^{-1}$ 1310 kJ

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web 6 a 218 g sample of steam at 121 $^\circ\text{C}$ is cooled to ice at 14 $^\circ\text{C}$ find the change in heat content of the system 7 in going from ice at 34 $^\circ\text{C}$ to steam at 138 $^\circ\text{C}$ a sample of water absorbs $1.41 \times 10^5 \text{ J}$ find the mass of the sample answers 1 2 68 x 104 J 2 6 23 x 104 J 3 1 71 x 104 J 4 1 11 x 105 J 5 2 11 x 105 J 6

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