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apr 25 2017 this is understood when you consider the two chemicals appear unchanged before and after the reaction



therefore the net ionic equation is  $2\text{Cl}^- + 2\text{Ag}^+$

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ions that remain essentially unchanged during a reaction are called spectator ions this means you can ignore them when you write the ionic equation you only need to model how the solid

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$1\text{K}^+ + \text{PO}_4^{3-} + \text{Al}^{3+} + \text{OH}^- \rightarrow \text{AlPO}_4 + 3\text{H}_2\text{O}$  ionic  $3\text{Cl}^- + \text{Al}^{3+} \rightarrow \text{AlCl}_3$   
 $3\text{Na}^+ + \text{PO}_4^{3-} + 3\text{K}^+ + \text{AlPO}_4 \rightarrow 3\text{K}^+ + \text{AlPO}_4 + 3\text{Na}^+$   
 $\text{FeCl}_3 + 3\text{Mg} \rightarrow \text{Fe} + 3\text{MgCl}_2$

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oct 2 2019 the net ionic equation is a chemical equation for a reaction that lists only those species participating in the

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aug 10 2023 write the full ionic and net ionic equations for this reaction solution the elemental metals magnesium on the reactant side copper on the product side are neutral

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the first step to writing a net ionic equation is balancing the chemical equation present let s use the reaction between sodium chloride and silver nitrate as an example next we write the

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highlights learning objectives by the end of this section you will be able to derive chemical equations from narrative descriptions of chemical reactions write and balance chemical

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the net ionic equation is the chemical equation that shows only those elements compounds and ions that are directly involved in the chemical reaction notice that in writing the net ionic

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what remains when the spectator ions are removed is called the net ionic equation which represents the actual chemical change occurring between the ionic compounds  $\text{Cl}^-$  and  $\text{Ag}^+$   
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 this net ionic equation indicates that solid silver chloride may be produced from dissolved chloride and silver ions regardless of the source of these ions these molecular and **net ionic equations practice khan academy** - Jul 18 2023  
 write the overall chemical equation the complete ionic equation and the net ionic equation for the reaction of aqueous silver fluoride with aqueous sodium phosphate to give solid silver

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the complete ionic equation for this reaction is as follows  

$$4\text{NH}_4\text{Cl} + 3\text{Ag}_2\text{CrO}_7 \rightarrow 12\text{AgCl} + 2\text{Cr}_2\text{O}_7^{2-} + 4\text{NH}_4^+$$
  
 because two  $\text{NH}_4^+$  are required to balance the charge of the  $\text{Cr}_2\text{O}_7^{2-}$  ion  
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