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 the general form  $ce_{aq} + h_{aq} \rightleftharpoons a_{aq}$  label 7 the equilibrium  
 constant expression for this reaction is  $K = \frac{ce_{aq} h_{aq}}{a_{aq}}$   
 label 8 where we denote the equilibrium constant  $K$  with a  
 subscript a for acid

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 investigating using the information provided by the  
 introduction within the virtual lab activity equilibrium  
 reaction  $co + h_2o \rightleftharpoons 62 + 4cl + cocl_2 + 6h_2o$  select each control test  
 tube to compare the concentrations of

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 chemical equilibrium in this semester you will also study  
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 learn about acids bases and salts unit 15 is a brief study of  
 the atomic nucleus which your teacher may introduce at any  
 time during the year

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web equilibrium moles  $hscn \cdot 2.00 \times 10^{-5} \text{ mol}$   $3.00 \times 10^{-6} \text{ mol}$   
 $1.70 \times 10^{-5} \text{ mol}$   $hscn$  knowing the number of moles of  $Fe^{3+}$  and  
 $hscn$  present in the equilibrium mixture and the volume of  
 the mixture we can easily find the concentrations of those  
 two species  $mol\ Fe^{3+} \cdot 1.70 \times 10^{-5} \text{ mol}$   $Fe^{3+} \cdot 8.50 \times 10^{-4} \text{ M}$   
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this unit out with a broaaaaad overview of what equilibrium is  
all about in chemistry we define equilibrium as the point at  
which the rate of the forward reaction is equal to the rate of  
the reverse reaction we learned about rates in unit five as  
how quickly a reaction proceeds

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web equilibrium questions a covalent bond has a standard  
gibbs free energy of 400 kJ/mol while the hydrogen bond and  
van der Waals forces have substantially lower values  
respectively 25 and 5 kJ/mol given such values which of the  
following is equivalent to the ratio of the  $K_{eq}$  value of a  
covalent bond to that of a hydrogen bond

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what is meant by chemical equilibrium two methods to  
describe the likely result of a chemical reaction in the  
previous chapter we discussed the principles of chemical  
kinetics which deal with the rate of change or how quickly a  
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reaction system what is its practical significance state the  
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explain the meaning of the statement equilibrium is  
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reactions occur at equal rates and the concentrations of  
products and reactants remain constant if we run a reaction  
in a closed system so that the products cannot escape we  
often find the reaction does not give a 100% yield of products  
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