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## A Gutmann

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web dec 13 2017 step 1 draw the lewis structure there are three central atoms  $\text{P}$   $\text{P}$   $\text{O}$  arrange them symmetrically in the order  $\text{P}$   $\text{O}$   $\text{P}$  attach the remaining  $\text{O}$  atoms as terminal atoms and arrange electrons to give each atom an octet the  $\text{P}$  atoms will get

expanded octets step 2 use vsepr theory to determine the electron geometries

9 2 lewis electron dot diagrams

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7 valence electrons chlorine has 7 valence electrons total valence electrons 1 iodine 7 3 chlorine 3 x 7 total valence electrons 7 21 total valence electrons 28 step 2 find the number of electrons needed to make the atoms happy iodine needs 8 valence electrons

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structure is basically a graphic  
representation of the electron

distribution around an atom the major  
reason why learning lewis dot structure  
is important is that it helps in  
predicting the number and type of

bonds which can be formed around an  
atom it also helps in predicting the  
geometry of the molecule