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Ensheng Dong

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substances ce a and ce b the mole
fractions of each would be written as
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c 6 h 14 d 0 660 g cm^3 when it is
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 cm^3 in a 35 0 55 0 ml ratio assume that
the volumes are additive
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web notice that the mole fraction has
no units on it and is written as a
decimal value do not change it to
percent note of caution you could see
the term mole percent it is simply the
mole fraction multiplied by 100 for
example in the problem just below the
mole fraction of cinnamic acid is 0 2885
its mole percent would be 28 85
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g/ml and xylene C_8H_{10} has a density
of 0.864 g/ml a solution was prepared
by mixing acetone and xylene calculate
the mole fraction of acetone in a
solution if 19.0 ml of acetone was
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unweighed quantity of O_2 are mixed in
a flask at room temperature