

# The 5 Functions Of Management

**Roman Wölfel**

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this goal was to give us experience finding the

standardization of through the use of primary standard

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titration the concentration of the acid or base is determined

by accurately measuring the volumes used in the

neutralisation reaction an indicator can be added to show the

end point of the reaction chemistry in society chemical

analysis analytical methods

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web introduction this laboratory exercise relies on a titration

technique to determine an unknown concentration of

monoprotic acid in solution in the process of titration a basic solution is gradually added to the acidic solution until

complete neutralization is obtained

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web with a strong acid and strong base titration strong acids

and strong bases refer to species that completely dissociate

to form ions in solution by contrast weak acids and bases

ionize only partially and the ionization reaction is reversible

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titration process is used in an acid base experiment in order

to determine the concentrations of solutions of acids and

bases through the titration process we are able to identify

physical changes to the mixture such as the colour change to

indicate the end point of the experiment

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web aug 30 2022 acid base titrations are usually used to

find the amount of a known acidic or basic substance

through acid base reactions the analyte titrand is the

solution with an unknown molarity the reagent titrant is the

solution with a known molarity that will react with the

analyte

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web conclusion the purpose of the experiment was to identify

the unknown solid acid 17 by calculating its molar mass this

was done by titration of the acid with sodium hydroxide in

order to find the volume of sodium hydroxide required to

reach the end point and that value was used in unit analysis to find the molar mass and identify the unknown solid acid

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web by utilizing the half equivalence point on the titration curve the pka of the acid analyte was determined to be 3.655 furthermore using the stoichiometric ratio between the naoh and the acid analyte 1:1 we could experimentally determine the number of moles of the acid in the solution to be 0.00731 moles making the experimental molar mass

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web chem 1210 spring 2019 experiment 10.11 part 1 acid base titration abstract the purpose of this experiment is to observe the titration of hydrochloric acid a strong acid with sodium hydroxide a strong base and acetic acid a

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web goals understand that parts of an acid base titration be able to determine the  $K_a$  or  $K_b$  from pH data associated with the titration of a weak acid or base be able to determine the molar mass of a solid monoprotic acid from titration data be able to calculate  $K_{a1}$  and  $K_{a2}$  for a polyprotic acid by the end of this lab students should be able to

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web titration experiment in the neutralization of hydrochloric acid by sodium hydroxide the mole ratio of acid to base is 1:1  $HCl(aq) + NaOH(aq) \rightarrow NaCl(aq) + H_2O(l)$   $HCl(aq) + NaOH(aq) \rightarrow NaCl(aq) + H_2O(l)$  one mole of HCl would be fully neutralized by one mole of NaOH

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web conclusion the identification of an unknown solid acid

the purpose of this experiment was to determine the standardization of sodium hydroxide NaOH and oxalic acid  $H_2C_2O_4 \cdot 2H_2O$  and the molar mass of an unknown solid acid using the molar mass of the standardized sodium hydroxide solution

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web acid base reactions can be used to determine concentrations of an acid or a base using stoichiometric calculations and titrimetric analysis the volume of a solution required to react

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web jul 19 2023 introduction titration is an analytical quantitative technique used to determine the concentration of a solute a pH titration is used to determine the concentration of an acid or a base titrations play an important role in determining amount and purity in many manufacturing processes

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web specifically an acid base titration can be used to figure out the concentration of an acid or base whether an unknown acid or base is strong or weak  $pK_a$  of an unknown acid or  $pK_b$  of the unknown base key terms titration a process where a solution of known strength is added to a certain volume of a treated sample containing an

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web conclusion this lab evaluated an acid base reaction with the volumetric analysis of a standard solution being NaOH reaction with a 2nd substance  $H_2C_2O_4(aq) + 2NaOH(aq) \rightarrow Na_2C_2O_4$

4h20 to differentiate an acid from a base in a solution a ph scale is used the ph scale ranges from 0 to 14 a ph of 7 is neutral in an acid base reaction

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web conclusion in this lab we used titration to explore the concepts of stoichiometry and equivalence points we found the concentration of an unknown substance by mixing 2m hcl with the naoh of unknown concentration in order to experimentally ascertain the concentration of the naoh

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web this point of titration is referred to as the equivalence point considering the 1 1 stoichiometry of this acid base reaction  $\text{NaOH} + \text{C}_6\text{H}_4\text{COOH} \rightarrow \text{NaC}_6\text{H}_4\text{COO} + \text{H}_2\text{O}$  the point of equivalence is the point of titration when the number of moles of naoh na added is equal to the number of moles of khp nb in the

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web astm a923 is designed to detect sigma phase in 22 cr duplex and superduplex1 it does not address nitrides or alpha prime later a second standard was written astm 1084 for lean duplex it quickly became apparent that astm a923 had some serious drawbacks

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