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RJ Shavelson

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benzoic acid is a versatile reagent that

can be implemented in a plethora of

syntheses e g benzene ref 1 methyl

benzoate ref 2 etc today benzoic acid is

most commonly encountered as

benzoate

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usually a phase transfer catalyst is used

but as it can be seen it still works

without one patreon h

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the products of its reaction with i

kmno4 in aqueous acid and ii kmno4 in

aqueous naoh predict the alkene give

the following products on oxidative

cleavage with kmno4 in acidic solution

draw the product formed when

cyclohexene undergoes a reaction with

cold dilute basic solution of kmno 4

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mechanism semi passive chemical

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in situ chemical oxidation or isco

schemes involve the addition of a

chemical oxidant such as potassium

permanganate kmno4 which destroys

chlorinated solvents like tce in a

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started our studies by identifying a

theory level that properly describes the

competition between the pathways a d

presented in fig 1 fig 1 possible

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b c 1 c 2 c c 2 c 3 d c 3 c 4 full size

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powerful oxidant if you run it a high

temperature with concentrated acid

then phenylacetic acid loses co2 to give

toluene which is then oxidised to

benzoic acid same with styrene under

controlled conditions you can get the

diol push harder and that oxidises

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when treated with hot concentrated

acidic kmnox4 k m n o x 4 arenes are

oxidised to the corresponding

carboxylic acids for example toluene is

oxidised to benzoic acid

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$$\text{MnO}_4^- + 7\text{H}^+ + 3\text{e}^- \rightarrow \text{MnO}_2 + 4\text{H}_2\text{O}$$

$$\text{MnO}_4^- + 4\text{H}^+ + 3\text{e}^- \rightarrow \text{MnO}_2 + 2\text{H}_2\text{O}$$

$$\text{MnO}_4^- + 2\text{H}^+ + 3\text{e}^- \rightarrow \text{MnO}_2 + \text{H}_2\text{O}$$

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 examples

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