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get q water 4 18 j g c x 110 g x 26 6 c
25 0 c q water 550 j Δh q water 550 j

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the heart of all calorimetry problems
and calculations because the heat
released or absorbed at constant
pressure is equal to Δh the relationship
between heat and Δh_{rxn} is $\Delta h_{rxn} q_{rxn}$
 $q_{calorimeter} mcs\Delta t$ 7 3 14 7 3 14 Δh_{rxn}
 $x n q_{rxn} q_{calorimeter} mcs\Delta t$

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answer the following questions 1 if i
have 125 grams of water at a
temperature of 25 0 c and heat it to
145 0 c what will the change in enthalpy

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temperature vs time for the process in
1 3 i have 18 7 grams of an unknown
liquid chemical at a temperature of 350
c

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ml of 1 00 m hcl at 25 0 c is added to
25 0 ml of 1 00 m naoh at 25 0 c in a
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and the temperature rises to 32 0 c
calculate the enthalpy change in kj for
this reaction assume the densities of
the solutions are 1 00 g ml and that
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of the copper

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of the calorimeter and its contents
changes from 19 77 c to 33 90 c
calculate e and h for the following
reaction as written 2 nh₃ g 3 o₂ g 2 ch

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web choose 1 answer the final temperature is closer to t_1 than to t_2 a the final temperature is closer to t_1 than to t_2 the final temperature is exactly halfway between t_1 and t_2 b the final temperature is exactly halfway between t_1 and t_2 the final temperature is closer to t_2 than to t_1 c

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ms t h 5.650 grams H_2O 4.184 J/g °C 55.4 °C 1310 kJ

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web 6 a 218 g sample of steam at 121 °C is cooled to ice at 14 °C find the change in heat content of the system 7 in going from ice at 34 °C to steam at 138 °C a sample of water absorbs 1.41 x 10⁵ J find the mass of the sample answers 1 2 68 x 10⁴ J 2 6.23 x 10⁴ J 3 1.71 x 10⁴ J 4 1.11 x 10⁵ J 5 2.11 x 10⁵ J 6

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