

# Nati Jota En Tanga

**Michal Rosen-Zvi**

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web now realizing that heating and cooling curves represent the amount of heat absorbed or released by a substance

during phase changes remember heat uses the variable  $q$  here we have a heating curve versus a cooling curve in a heating curve our substance is absorbing heat

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web nov 15 2021 for the heating and cooling curve of any given substance the solid phase of that substance will be represented in the lower left corner of the graph where the temperature is at its lowest

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 isolated from its environment is subjected to heat changes  
 corresponding changes in temperature and phase of the  
 substance is observed this is graphically represented by  
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web determine the final temperature and state of the  
 ammonia once the system reaches equilibrium assuming that  
 no heat is lost to the surroundings use the information below  
 normal boiling point of  $\text{NH}_3$  33.3 °C  $\Delta H_{\text{vap}}$  23.4 kJ/mol  
 specific heat  $\text{NH}_3$  1.474 J/g K specific heat  $\text{NH}_3$  2.06 J/g K  
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web apr 3 2023 the experiment described above can be  
 summarized in a graph called a heating curve figure below  
 figure 11 16 1 11 16 1 in the heating curve of water the  
 temperature is shown as heat is continually added changes  
 of state occur during plateaus because the temperature is  
 constant

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 block of ice that is at a temperature of 30 °C well below its  
 melting point the ice is in a closed container as heat is  
 steadily added to the ice block the water molecules will  
 begin to vibrate faster and faster as they absorb kinetic  
 energy

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 time in figure pageindex 4 plots temperature versus time as  
 a 75 g sample of steam initially at 1 atm and 200 °C is cooled  
 although we might expect the cooling curve to be the mirror  
 image of the heating curve in figure pageindex 3 the cooling  
 curve is not an

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web once all of the material has solidified at the time  
 indicated by point c the cooling will continue at a rate



determined by the heat capacities of the two solids a and b the composition and of course the geometry of the experimental set up by measuring cooling curves for samples of varying composition one can map the entire phase heating and cooling curves example 1 channels for pearson - Feb 27 2023

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web for water this temperature is 100 c because the boiling point for water is 100 c different substances have different melting points and boiling points but the shapes of their heating curves are very similar for example this is the heating curve for iron a metal that melts at 1538 c and boils at 2861 c cooling curves

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