

Oh Happy Day Text

Wolfgang Guggemos

What are the spectator ions in the reaction between sodium Oct 16, 2016 · The nitrate and the sodium ions.

$\text{Na}_2\text{CO}_3(\text{aq}) + 2\text{AgNO}_3(\text{aq}) \rightarrow \text{Ag}_2\text{CO}_3(\text{s}) + 2\text{NaNO}_3(\text{aq})$ The net ionic equation is: $2\text{Ag}^+(\text{aq}) + \text{CO}_3^{2-}(\text{aq}) \rightarrow \text{Ag}_2\text{CO}_3(\text{s})$...

What is the mass percent of hydrogen for C₃H₅OH? | Socratic Jan 2, 2017 · In order to compute

the mass percent of hydrogen in C₃H₅OH we're going to need two things: 1: The formula mass of the entire compound 2:

The atomic mass of hydrogen The formula mass of C₃H₅OH is 58.08 g/mol because: $3(12.01 \text{ g/mol}) + 6(1.008 \text{ g/mol}) +$

$(16.00 \text{ g/mol}) = 58.08 \text{ g/mol}$ The atomic mass of the hydrogen atom is 1.008 g/mol Now, we have to use the following

equation: The numerator ... **Question #627ef - Socratic** functional groups like -OH, -CHO, -CO- are present in

carbohydrates. commonly two functional are present in a carbohydrate, -OH and -CHO, or -OH and -CO-. Aldose shows

different reactions and test like Tollen's silver mirror test, Fehling's test. different types of reactions are also shown by

carbohydrates like Osazone synthesis, length increasing reaction, ... *0.10 moles of sodium cyanate dissolved in 250 cm³*

distilled Dec 28, 2015 · The cyanate anions will react with water to form isocyanic acid, HNCO, and hydroxide anions, OH⁻.

This tells you that you can expect the pH of the resulting solution to be higher than 7. Calculate the molarity of the sodium

cyanate solution by using the given number of moles and volume. Keep in mind that you need to use liters of solution and ...

Could a buffered solution be made by mixing aqueous ... - Socratic Mar 25, 2017 · Could a buffered solution be made

by mixing aqueous solutions of HCl and NaOH? Explain. Why isn't a mixture of a strong acid and its conjugate base

considered a buffered solution? **What is the formula of "magnesium hydroxide"? | Socratic** Feb 27, 2017 · $\text{Mg}(\text{OH})_2$

Now we know that hydroxides are salts of HO⁻, and some metal ion. Now if the parent metal has an electronic configuration

of 2:8:2, then there are 12 electrons, and the atomic number of the metal is equal to 12. We look on the Periodic Table, and

we find that Z=12, for "magnesium metal". As a Group 2 metal, magnesium forms a Mg^{2+} ion, and hence its hydroxide is

$\text{Mg}(\text{OH})_2$. **Question #44ec4 - Socratic** That's a silly way to say carbon dioxide! $\text{CO}_2(\text{g}) + \text{Ca}(\text{OH})_2(\text{aq}) \rightarrow \text{Ca}^{2+}(\text{aq}) + \text{H}_2\text{CO}_3(\text{aq})$

May be wrong, but it seems reasonable. The gas would interact with the hydroxide and water

molecules to form carbonic acid. There are also concentrations of OH⁻ and H⁺ leftover from this reaction and from the

auto-ionization of water, but that's probably going too far into detail. What is the product of the following reaction? 1)CH₃

OH - Socratic Dec 23, 2017 · These are ostensibly acid-base reactions.... For ammonium we could write.... $\text{NH}_4^+ + \text{HO}^-$

$\text{rarr NH}_3(\text{aq}) + \text{H}_2\text{O}(\text{l})$ For methanol, the acid base reaction would proceed ... *What is the chemical formula for lithium hydroxide?* - Socratic Apr 15, 2017 · Lithium is a Group 1 metal and commonly forms a M^+ ion. Hydroxide anion, OH^- , has a unit negative charge. When they make music together, there is thus 1:1 stoichiometry between ions: $\text{Li}(\text{s}) + \text{H}_2\text{O}(\text{l}) \rightarrow \text{LiOH}(\text{aq}) + \frac{1}{2} \text{H}_2(\text{g})$ ↑ □ □ **If the pH=9 for a solution of KOH(aq), what is the ...** - Socratic Aug 18, 2016 · If the pH = 9 for a solution of $\text{KOH}(\text{aq})$, what is the concentration of H_2O ?

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